



Europäisches Patentamt
European Patent Office
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⑪ Publication number:

0 181 078
A1

⑫

EUROPEAN PATENT APPLICATION

⑬ Application number: 85306885.8

⑬ Int. Cl.4: C 07 C 53/02
C 07 C 51/15

⑭ Date of filing: 26.09.85

⑮ Priority: 29.09.84 GB 8424672

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⑯ Date of publication of application:
14.05.86 Bulletin 86/20

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⑯ Designated Contracting States:
AT BE CH DE FR GB IT LI NL SE

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⑲ Process for the preparation of formic acid.

⑳ An integrated process for the production of formic acid from carbon dioxide and hydrogen is provided. The process comprises for example, a reactor in which the formate salt of a nitrogenous base is produced together with subsequent purification and formic acid recovery stages. In the process described formic acid is recovered from the formate salt by using a base interchange reaction in which there is generated a thermally decomposable formate salt of a high boiling base. The high boiling base can be for example an imidazole and the nitrogenous base can be triethylamine.

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PROCESS FOR THE PREPARATION OF FORMIC ACID

The present invention provides a process for the production of formic acid from carbon dioxide and hydrogen using an integrated series of process steps.

Our European patent applications 0095321 and 84301772.4 5 respectively describe a method for the production of a trialkylammonium formate from a tertiary amine, carbon dioxide and hydrogen and a method for converting the trialkylammonium formate into another formate salt which is thermally decomposable to formic acid.

10 An integrated process has now been devised which allows formic acid to be prepared from carbon dioxide and hydrogen feedstocks only.

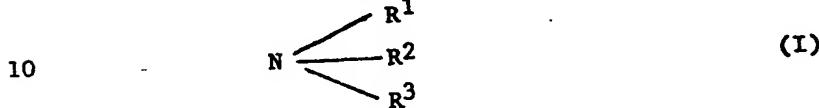
Accordingly, the present invention provides an integrated process for the production of formic acid from carbon dioxide and 15 hydrogen characterised in that

- (a) in a first stage a nitrogenous base, carbon dioxide and hydrogen are reacted together in the presence of a catalyst to produce a formate salt of the nitrogenous base.
- (b) in a second stage the catalyst is removed from the formate salt of the nitrogenous base and any low boilers and recycled to the 20 first stage;
- (c) in a third stage the formate salt of the nitrogenous base is recovered from the low boilers.
- (d) in a fourth stage the formate salt of the nitrogenous base is 25 reacted with a base having a high boiling point to produce the

nitrogenous base and the formate salt of the base having a high boiling point.

(e) in a fifth stage the formate salt of the base having a high boiling point is decomposed to the higher boiling base and formic acid.

5 By the term nitrogenous base is meant a nitrogenous base containing a tertiary nitrogen atom. The nitrogenous base containing a tertiary nitrogen atom may suitably be of formula



or of formula:



wherein in the formulae, R¹, R² and R³, which may be the same or different, are hydrocarbyl groups or substituted hydrocarbyl groups or any two or all of R¹, R² and R³ may form part of a ring, R⁴ is a hydrocarbyl group or substituted hydrocarbyl group and R⁵ is a divalent organic group or R⁴ and R⁵ may form part of a ring. Suitably the hydrocarbyl group is an aliphatic, cycloaliphatic, aryl, or alkaryl group. Substituted hydrocarbyl groups may contain for example nitrogen or oxygen. Preferably the organic base is a trialkylamine, even more preferably a lower trialkylamine, for example a C₁ to C₁₀ trialkylamine. Examples of suitable trialkylamines are trimethylamine, triethylamine, tripropylamine and tributylamine. Examples of other nitrogenous bases which may be employed are 1,8-diazabicyclo[5.4.0]undec-7-ene (DBU) and 1,4-diazabicyclo[2.2.2]octane (DABCO), pyridines and picolines. Mixtures of nitrogenous bases may also be used.

30 The formate salt produced in the first stage of the process corresponds to the nitrogenous base used in this stage. Thus using triethylamine the product is triethylammonium formate.

As regards the base having a high boiling point, this is suitably also a nitrogenous base and is selected so that

35 (1) it is weaker than the nitrogenous base used in the first stage

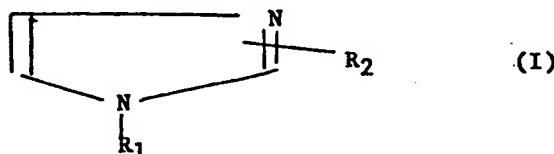
of the process.

(2) its formate salt is thermally decomposable at a temperature higher than the boiling point of the nitrogenous base used in the first stage of the process, and

5 (3) it is less volatile than the nitrogenous base used in the first stage of the process.

It will be seen from these criteria that the exact choice of such a base will depend upon which nitrogenous base is used in the first stage of the process.

10 Preferably the second base has a pK_a in the range 4.0 to 9.0 and is an imidazole of the general formula:



15 where R_1 is a monovalent hydrocarbon group containing 1 to 12 carbon atoms and R_2 is a hydrogen atom or an R_1 group, the total number of carbon atoms and R_1 and R_2 conveniently being not more than 20 and preferably from 4 to 12.

20 Suitable hydrocarbon radicals on the imidazole derivatives (I) are, in general, alkyl groups of 1 to 8 carbon atoms, cyclopentyl, cyclohexyl, phenyl and methylphenyl groups. Amongst the above, imidazole derivatives where R_1 is *n*-1-alkyl of 4 to 10 carbon atoms and R_2 is hydrogen or methyl are particularly suitable.

25 Examples of such compounds are 1-(*n*-1-butyl)-imidazole (pK_a 5.9), 1-(*n*-1-pentyl)-imidazole (pK_a 5.9), 1-(*n*-1-decyl)-imidazole (pK_a 5.75), 1-(*n*-1-butyl)-2-methylimidazole (pK_a 7.0) and 1-(*n*-1-pentyl)-2-methylimidazole (pK_a 6.85).

30 In addition to imidazoles, quinoline and other heterocyclic nitrogenous bases can be used.

For a definition of the pK_a values, which are a measure of the base strength, reference may be made, for example to Landoldt-Bornstein, 6th edition, 7th part volume II, page 900 et seq.

35 As catalyst there is used a compound of a Group VIII transition

metal, which is either iron, nickel, ruthenium, rhodium, palladium, iridium or platinum. Suitably the metal is ruthenium or rhodium and is preferably ruthenium. Mixtures of compounds of different transition metals may also be used if so desired. The metal or metals may be added in any convenient form which is soluble in the reaction mixture. Thus the metal or metals may be added in the form of a simple salt, for example a halide, or in the form of a complex, for example a hydride complex. Examples of suitable ruthenium compounds which may be employed as catalyst are $\text{RuCl}_2(\text{PPh}_3)_3$,
10 $\text{RuH}_2(\text{PPh}_3)_4$, $\text{RuHCl}(\text{PPh}_3)_4$, $\text{RuCl}_3 \cdot 3\text{H}_2\text{O}$, $[\text{Ru}(\text{CO})_2\text{Cl}_2]_n$, $[\text{Ru}(\text{CO})_2\text{I}_2]_2$,
[p-cymene] $\text{RuCl}_2]_2$, [(hexamethylbenzene) $\text{RuCl}_2]_2$ and
[(hexamethylbenzene) $\text{Ru}_2(\text{OH})_3\text{Cl}$. Suitably the catalyst concentration may be in the range 50 to 5000, preferably from 250 to 1000 parts per million by weight (as the metal in the complex).

15 The operation of the process is now described with reference to the Figure, which illustrates a schematic example of a plant using the process.

20 A recycle stream containing the nitrogenous base and water is fed via line (4) to the first stage, reactor (6), along with three other streams, (a) a hydrogen stream (via line (8)), (b) a carbon dioxide stream (via line (10)) and (c) a catalyst stream (via line (12)). The catalyst stream preferably comprises a solution of the catalyst in a high boiling point solvent.

25 It is important that the high boiling solvent is selected so that

- (i) the catalyst is soluble in the recycle stream (12) (ie the solvent acts as a catalyst carrier),
- (ii) the solvent does not adversely affect the activity of the catalyst in the reaction stage (6) and,
- 30 (iii) the nitrogenous base formate produced in the first stage can readily be separated from the solvent.

The high boiling solvent may be for example either one or more alcohols.

35 Examples of solvents include, diethylene glycol, tetraethylene glycol, polyethylene glycol, 1-phenyl-1-propanol and

3-phenyl-1-propanol. Sulpholanes can also be used. It is preferred to use as solvent either tetraethylene glycol or polyethylene glycol (average MW 400). In the reactor (6) it is preferable though not essential also to have some water present as this increases the reaction rate.

5 The reactor (6) is suitably operated at a temperature in the range from 20 to 200°, preferably from 60 to 130°C.

The carbon dioxide may either be carbon dioxide itself, which is widely available on an industrial scale, or a carbon dioxide 10 obtained from a carbonate or a bicarbonate sources. Alternatives carbonates or bicarbs can be used directly. Carbon dioxide may be used as a gas or as a liquid or as a solid. Using carbon dioxide gas as the source of carbon dioxide it is preferred to use partial pressures of carbon dioxide and hydrogen which are as high as is 15 practicable and economic. The use of high partial pressures of hydrogen is desirable because the reaction rate and yield of the formate salt increase as the partial pressure increases. The partial pressure of carbon dioxide is less critical but suitably the carbon dioxide partial pressure may be up to 60 bar and the hydrogen 20 partial pressure up to 250 bar.

Conveniently the partial pressure of carbon dioxide is from 10 to 50 bar and that of hydrogen from 10 to 150 bar. The ratio of the partial pressure of hydrogen to that of carbon dioxide in reactor (6) is preferably at least 1:1 more preferably at least 1.5:1.

25 The product from the first stage reactor comprises unreacted materials, the formate salt of the nitrogenous base and catalyst in a water/high boiling solvent mixture. This product from the first stage reactor (6) is fed via line (14) to a second stage (16) where the catalyst and the high boiling solvent are removed.

30 The second stage (16) may suitably comprise:

- (a) an evaporator, where (i) the catalyst and the high boiling solvent are separated and recycled to the first stage reactor (6) via line (12) and (ii) the gaseous components are separated and recycled via line (18) followed by
- 35 (b) a unit for the separation of unreacted nitrogenous base and

water (also for recycle via line (18)) from the formate salt of the nitrogenous base. This may take the form of (i) a decanter (to separate any aqueous and organic phases) or (ii) a distillation column. The formate salt of the nitrogenous base is removed from the second stage via line (22).

5 Alternatively the second stage (16) can be a distillation tower where the gaseous components i.e. carbon dioxide, hydrogen and the low boilers, i.e. the nitrogenous base and some water are removed overhead and the intermediate boilers i.e. the formate salt of the 10 nitrogenous base and some water are removed at an intermediate point in the column via line (22). The gaseous products and low boilers are recycled via line (18) while the non volatile catalyst/high boiling solvent mixtures recycled as a liquid to the first stage reactor (6) via line (12).

15 After the second stage, the (aqueous) solution of the formate salt of the nitrogenous base, is fed via line (22) to the third stage (24) which can consist for example of a reaction kettle with a distillation column mounted on top. The formate salt of the 20 nitrogenous base is fed to the column along with an appropriate base having a high boiling point fed via line (20). By maintaining the column at high temperature a base interchange reaction is set up with the result that the nitrogenous base is liberated and the formate salt of the base having a high boiling point is formed. The 25 nitrogenous base is removed overhead in the distillation column and recycled via line (26) thereby driving the base interchange reaction to completion. The product, which comprises a solution of the formate salt of the base having a high boiling point is removed from the kettle via line (28) and fed to the fourth and final stage (30).

The rate of the base interchange reaction in the third stage is 30 increased by using high temperatures e.g. one greater than the boiling point of the nitrogenous base and subatmospheric pressure. It is preferred therefore to use as high a temperature as possible, without producing undesirable decomposition products, and to use as low a distillation pressure as is economically 35 practicable.

In the fourth stage (30) the feed from line (28) is heated to a temperature at which the formate salt of the base having a high boiling point decomposes to formic acid and the base having a high boiling point. It is preferable to use subatmospheric pressure to reduce the formation of side products. The formic acid is removed overhead via line (32) and can be further purified if desired while the remaining solution of the high boiling base is recycled to line (22) via line (20).

The recycle components are fed via line (18) and line (26) to a gas scrubber (2). The gas scrubber is an optical feature. Inside the gas scrubber (2) the recycle components are contacted with carbon dioxide in the form of a gaseous stream fed via line (34) and vented via line (36). The gaseous stream of carbon dioxide can be, for example

- 15 (1) low value carbon dioxide from another plant such as a fermentation process or boiler flue.
- (2) a carbon dioxide containing off-gas from another part of the process,
- (3) carbon dioxide derived from a unit specifically designed for the manufacture of carbon dioxide.
- 20 (4) carbon dioxide produced by a plant which makes CO₂ as a byproduct.
- (5) carbon dioxide from a gas field.

In the gas scrubber (2), some or all of the carbon dioxide present in the gaseous stream is removed and solubilised in the recycle components.

During the solubilisation of the gas, the carbon dioxide reacts with the nitrogenous base and water present in the recycle components to produce at least in part, bicarbonate of the nitrogenous base. This bicarbonate provides a useful method of carrying carbon dioxide via line (4) to the first stage reactor (6) where it is also converted to the corresponding formate.

The invention is illustrated by the following Examples.

Example 1 - Production of Triethylammonium Formate

35 A liquid feed of composition 21.9% w/w triethylamine, 76.4% w/w

tetraethylene glycol with 1.7% w/w water containing ruthenium catalyst (500 parts per million in ruthenium) was introduced into a one litre stirred stainless steel reactor at a rate of 2000 ml/hour. Liquid carbon dioxide was also introduced at 5 232g/hour and hydrogen fed to keep the pressure of the reactor at 100 bar; a liquid product was taken from the reactor in order to keep the reactor inventory as near to 500g as possible. A 2.67 hour mass accountability study was carried out following a 5.0 hour prerun to establish steady operating conditions. During the balance 10 period, 5199g feed were introduced and 5774g of liquid product obtained. The reactor productivity was estimated as 6.2 moles/Kg reactor charge/hour. The reactor product contained triethylammonium formate (20.9% w/w as a 1:1 adduct).

Example 2 - Base Interchange Reaction

15 A feed of composition 63.4% w/w 1-n-butylimidazole, 34.8% w/w triethylammonium formate ($[\text{NEt}_3][\text{HCO}_2\text{H}]_2$), 1.8% w/w water was introduced into a distillation column (set for total take-off) at a kettle temperature of 178°C and atmospheric pressure. During a mass accountability study following a pre-run to achieve steady operating 20 conditions, 1940.9g of feed were introduced, and 858.2 and 160.7g of base and overhead products obtained. Analysis of the feed and products showed the following respective % w/w's of triethylamine (present in the feed as the formate) and 1-n-butylimidazole (present in the product as the formate). Feed 18.7 and 63.4, Base product 25 2.9 and 76.2 and Overhead product 91.0 and 0.1.

Example 3 - Base Formate Decomposition

30 A feed of composition 93.5% w/w 1-n-butylimidazolium formate, 4.7% w/w 1-n-butylimidazole and 1.8% w/w water was introduced into a distillation column at 185°C and 150 mm Hg. A reflux ratio of 0.5 was employed. During a mass accountability study of 1 hour, following a pre-run to achieve steady operating conditions, 436.6g 35 of feed were introduced and 382.0 and 71.7g of base and overhead products obtained. The base and overhead products contained respectively 17.5 w/w formic acid (present as 1-n-butylimidazolium formate) and 88.0% w/w formic acid.

Claims:

1. An integrated process for the production of formic acid from carbon dioxide and hydrogen characterised in that
 - (a) in a first stage a nitrogenous base, carbon dioxide and hydrogen are reacted together in the presence of a catalyst to produce a formate salt of the nitrogenous base.
 - (b) in a second stage the catalyst is removed from the formate salt of the nitrogenous base and any low boilers and recycled to the first stage;
 - (c) in a third stage the formate salt of the nitrogenous base is recovered from the low boilers;
 - (d) in a fourth stage the formate salt of the nitrogenous base is reacted with a base having a high boiling point to produce the nitrogenous base and the formate salt of the base having a high boiling point.
 - (e) in a fifth stage the formate salt of the base having a high boiling point is decomposed to the higher boiling base and formic acid.
2. An integrated process as claimed in Claim 1 characterised in that the first stage is carried out in the presence of a high boiling solvent and that in the second stage the catalyst is removed and recycled as a solution in the high boiling solvent.
3. An integrated process as claimed in Claim 2 characterised in that the high boiling solvent is selected from diethylene glycol, tetraethylene glycol, polyethylene glycol, 1-phenyl-1-propanol and 3-phenyl-1-propanol.

4. An integrated process as claimed in Claim 1 or Claim 2 characterised in that the catalyst is a ruthenium catalyst.

5. An integrated process as claimed in Claim 1 or Claim 2 characterised in that the nitrogenous base is a C₁ to C₁₀ trialkylamine.

6. An integrated process as claimed in Claim 1 or Claim 2 characterised in that the base having a high boiling point is an imidazole or a quinoline.

7. An integrated process as claimed in Claim 1 or Claim 2 10 characterised in that the fourth stage is carried out at temperature above the boiling point of the nitrogenous base and subatmospheric pressure.

8. An integrated process as claimed in Claim 1 or Claim 2 15 characterised in that the fifth stage is carried out at subatmospheric pressure.

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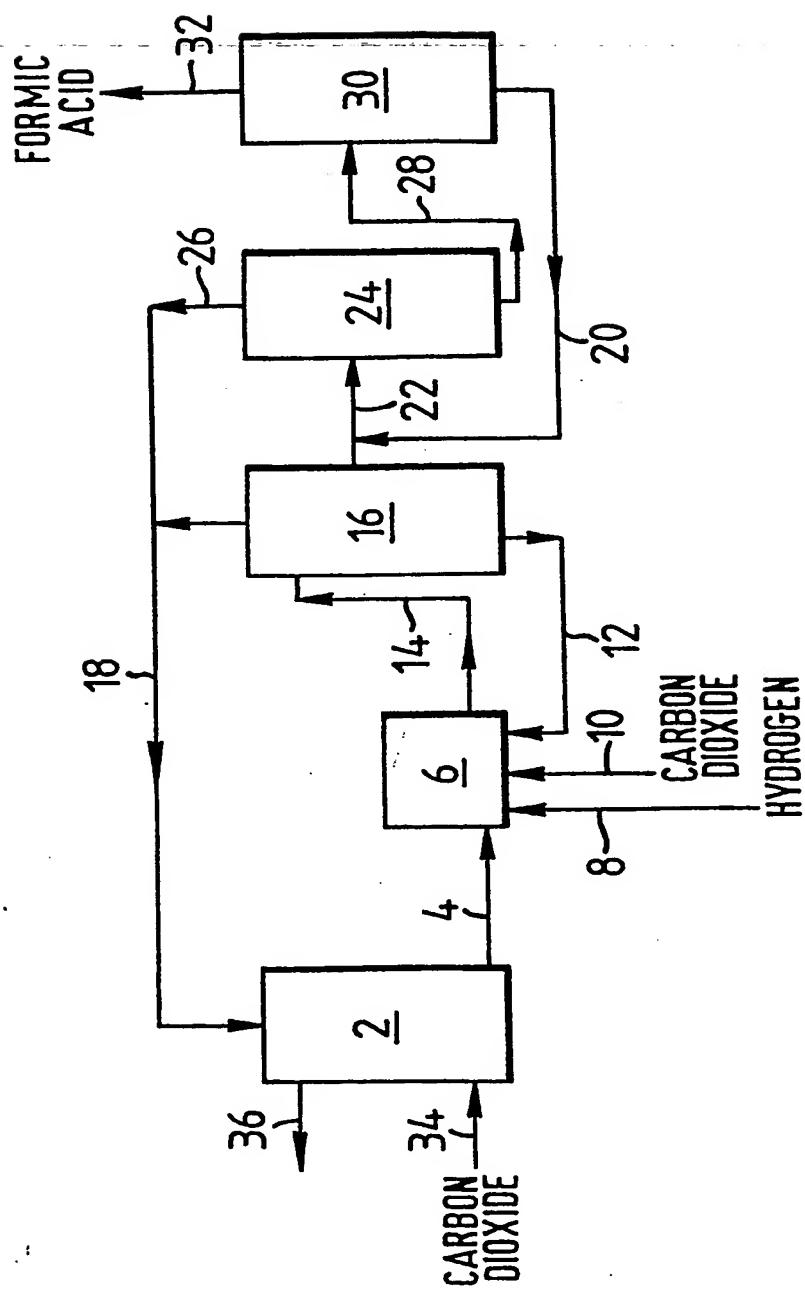
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DOCUMENTS CONSIDERED TO BE RELEVANT			
Category	Citation of document with indication, where appropriate, of relevant passages	Relevant to claim	CLASSIFICATION OF THE APPLICATION (Int. Cl 4)
Y,D	EP-A-0 095 321 (BP CHEMICALS LTD.) * claim 1 *	1	C 07 C 53/02 C 07 C 51/15
A,D	* claim 8 *	4	
Y,P D	EP-A-0 126 524 (BP CHEMICALS LTD.) * claim 1 *	1	
Y	US-A-3 414 610 (K. WAGNER et al.) * claims 1, 8 *	1	
A	* claims 5, 6 *	5	TECHNICAL FIELDS SEARCHED (Int. Cl.4)
	-----		C 07 C 51/02 C 07 C 51/15 C 07 C 53/02
The present search report has been drawn up for all claims			
Place of search	Date of completion of the search	Examiner	
BERLIN	19-12-1985	KNAACK M	
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